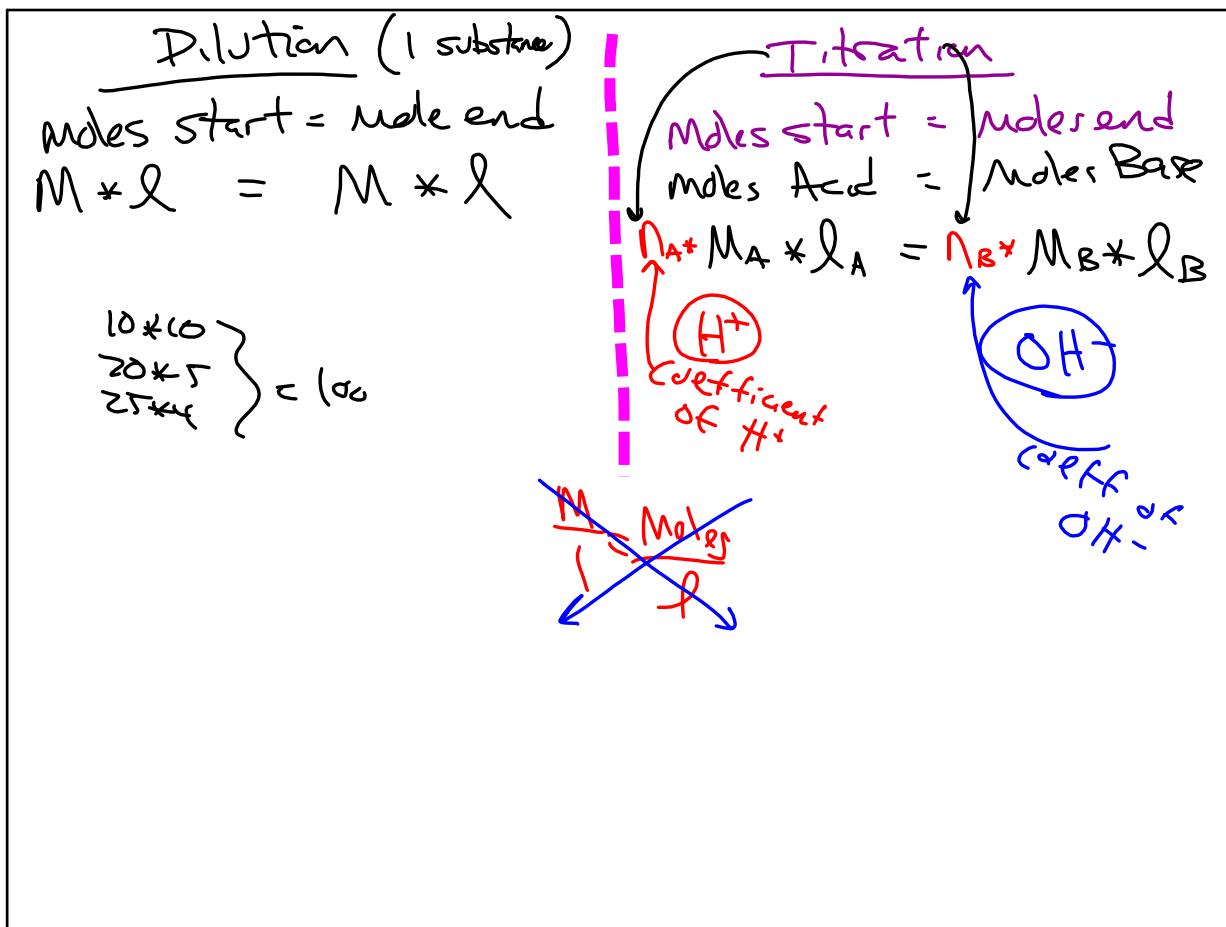


Oct 3-8:04 AM

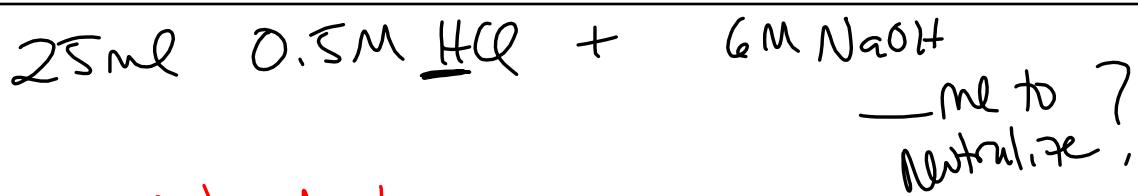


Oct 3-8:26 AM

Equivelance Pt  $\Rightarrow$  Moles Acid and Base equal  
Neutral

End point  $\Rightarrow$  indicator changes color  
Phenolphthalein Dyes at  $\approx$  pH = 9

Oct 3-8:34 AM



$$\text{Moles Acid} = \text{Moles Base}$$

$$1 \text{ M L} = 1 \text{ M L}$$

$$(1)(0.5)(25\text{ mL}) = (1)(6) \text{ mL}$$

$$2.08 \text{ mL}$$

Oct 3-8:36 AM

PS4

(20) \_\_\_\_\_ moles BaCl<sub>2</sub> (Nett) 393 ml 0.171 M Ba(OH)<sub>2</sub>

With HCl

2 HCl + Ba(OH)<sub>2</sub> → BaCl<sub>2</sub> + 2 H<sub>2</sub>O

393 ml  
0.171 M

Moles = M \* l  
= (0.171)(0.393)  
= 0.0672 moles Ba(OH)<sub>2</sub>

0.0672 moles | MOLE RATIO 0.1344 moles

Oct 3-8:41 AM

PS4 (20) \_\_\_\_\_ l 0.25 M HNO<sub>3</sub> (Nett) 17.5 g NaOH in 350 ml

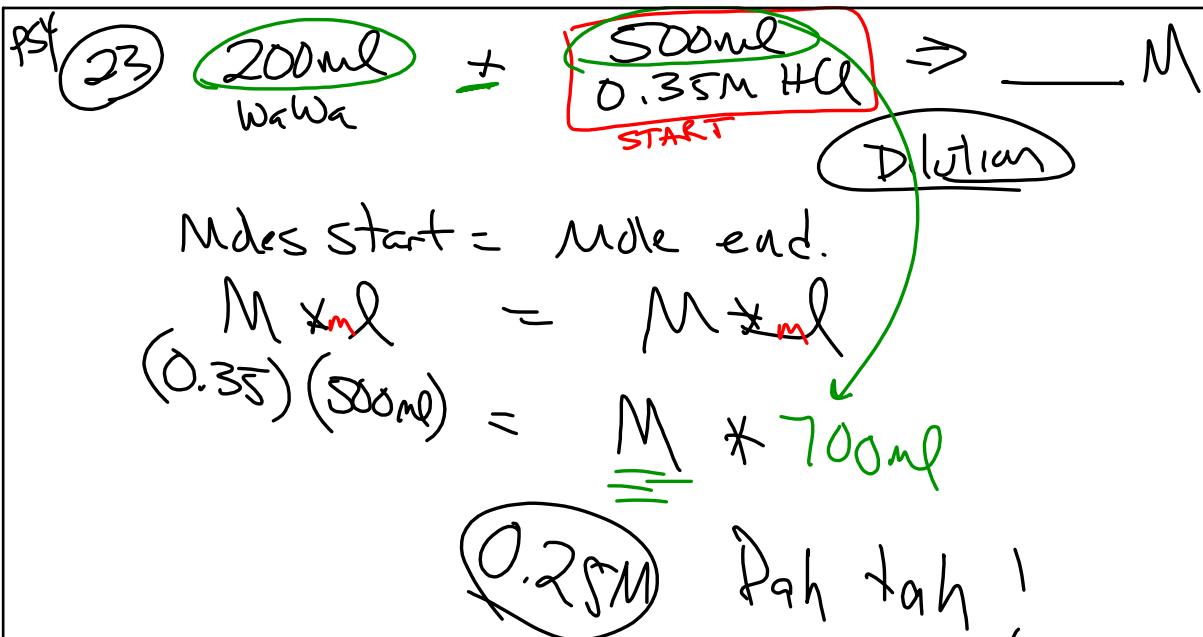
Moles Acid = Moles Base

l \* M \* l  
(1)(0.25) l = 0.4375

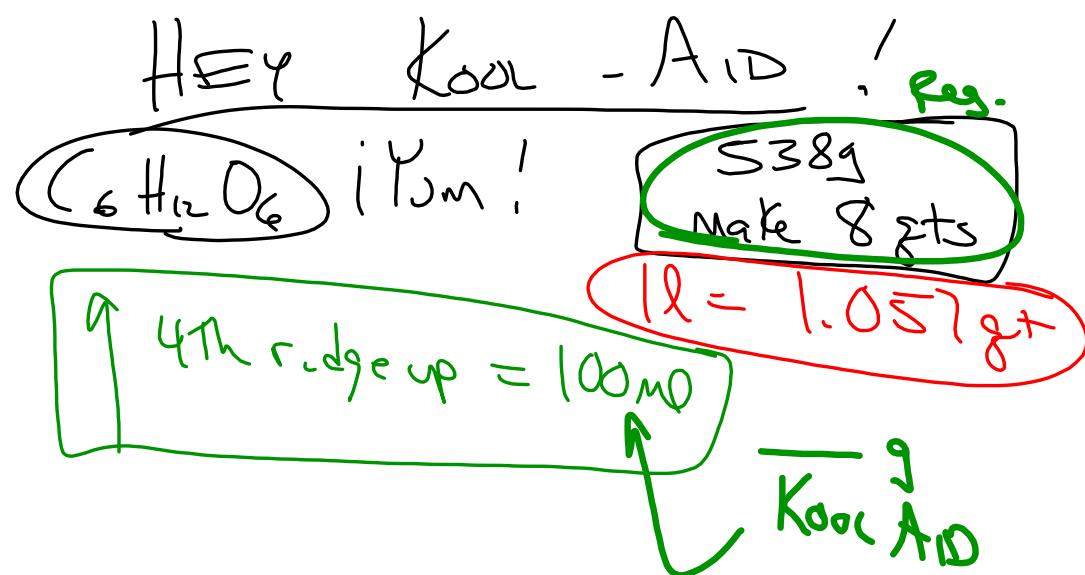
1.75 l

17.5 g NaOH / 1 mole NaOH = 40 g NaOH

Oct 3-8:50 AM



Oct 3-8:56 AM



Oct 3-9:01 AM

$$\frac{538 \text{ g C}_6\text{H}_{12}\text{O}_6}{8 \text{ g}} \times \frac{1.057 \text{ g/L}}{1000 \text{ mL}} = 7.11 \text{ g C}_6\text{H}_{12}\text{O}_6$$

HW - Finish Exam |  
and Ec (odd)

Oct 3-9:11 AM