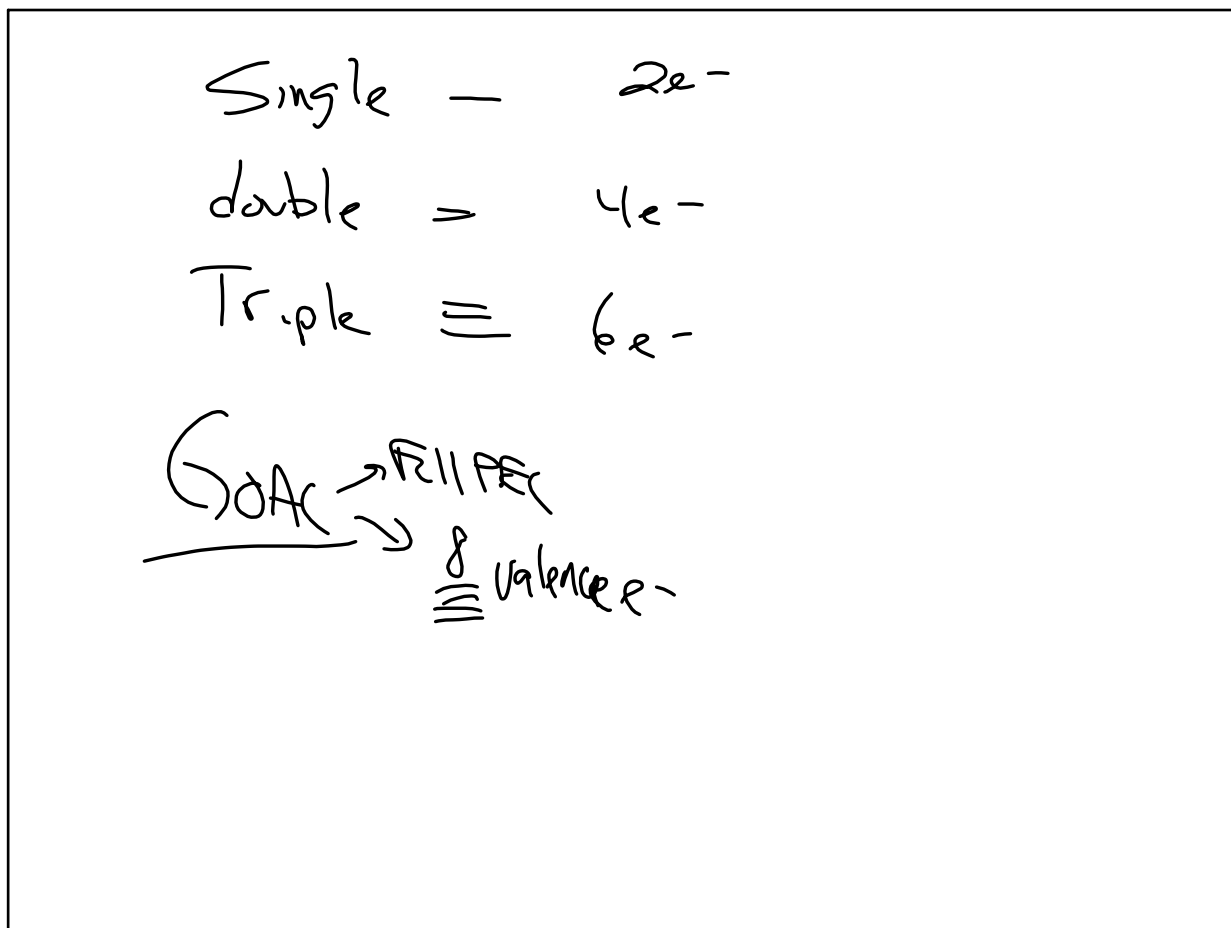


Nov 12-8:04 AM



Nov 12-8:43 AM

PCl_3 Lewis structure

Draw The bonds

$N-N$ 1.47 \AA

$N=N$ 1.24 \AA

$N \equiv N$ 1.10 \AA

more pull →
greater e density

Angstrom $\times 10^{-10} m$

2D → 3D

chap 8 chap 9

 P 347+351

 SSSS

Nov 12-8:45 AM

PCl_3 3 bond sites } ④
1 non-bond sites }

$5 + 3(7) = 26 e^-$

2D

① Total up all valence e- into the "Bank" (more metallic element)

② metal in middle (gives away)

③ Attach ^{outside} periphery to central ^{middle} atom with SB to start. Subl. # e used from bank.

④ Fix periphery → make stable

⑤ Fix central atom

If not enough e- to fix central atom

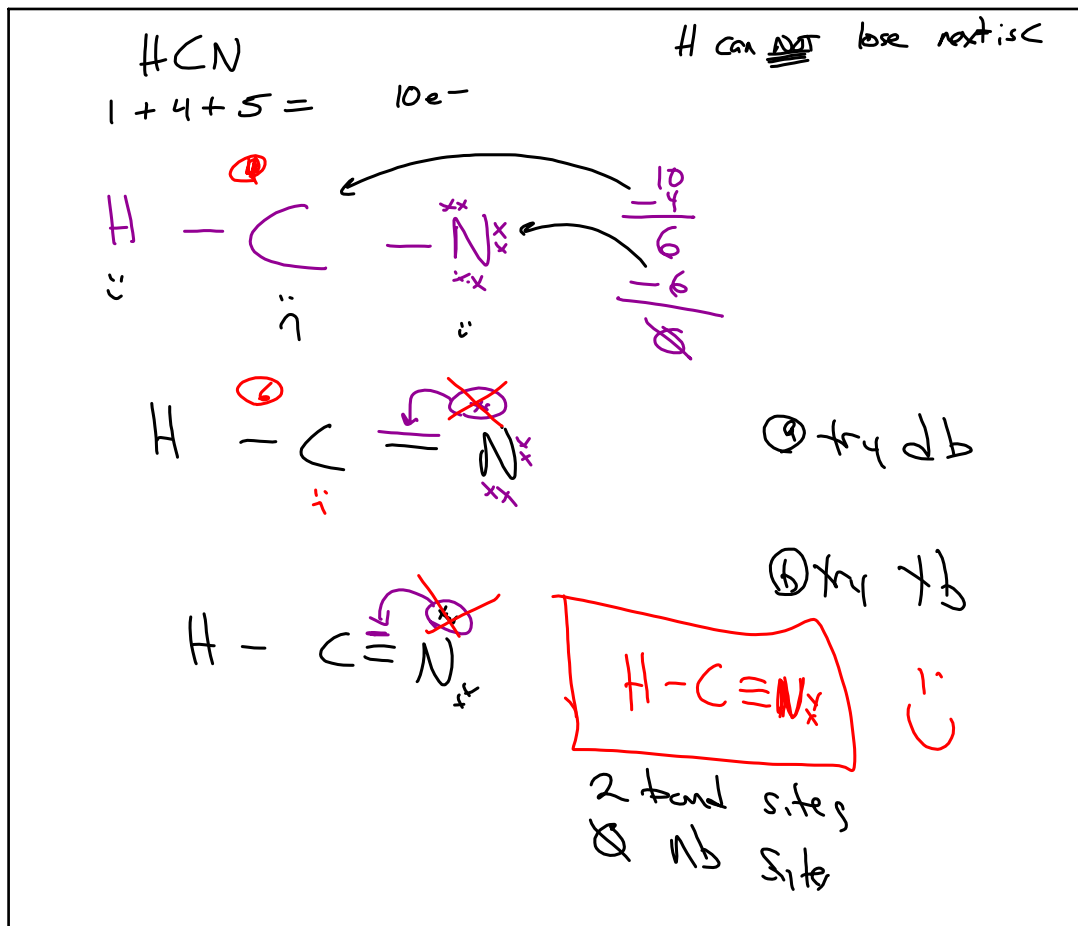
① try db ② try tb

26	
- 6	
20	
- 18	
2	
- 2	
0	

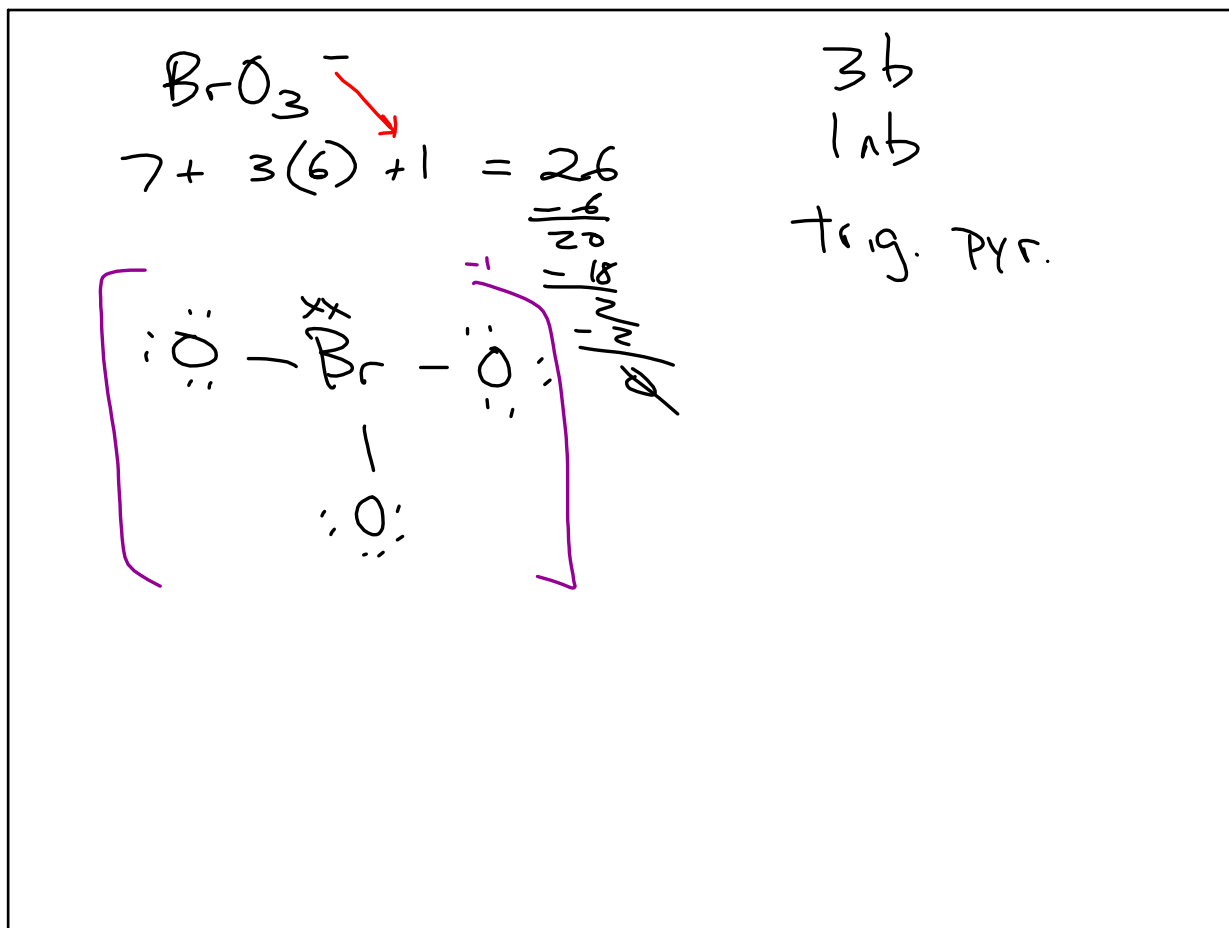
$P \leftrightarrow Cl$ Polar covalent between atoms

Molecule $PCl_3 \rightarrow$ Polar Molecule
 Not Symmetrical

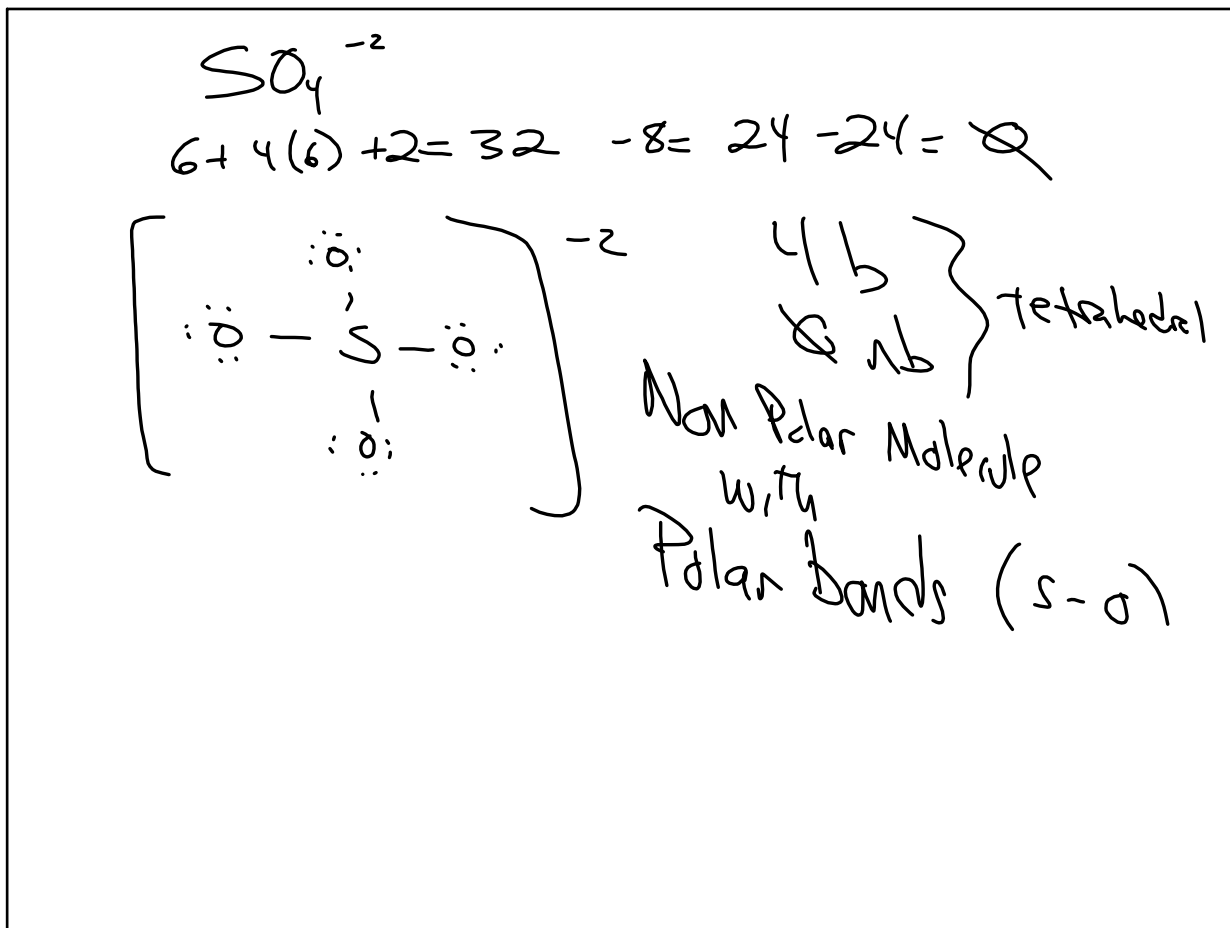
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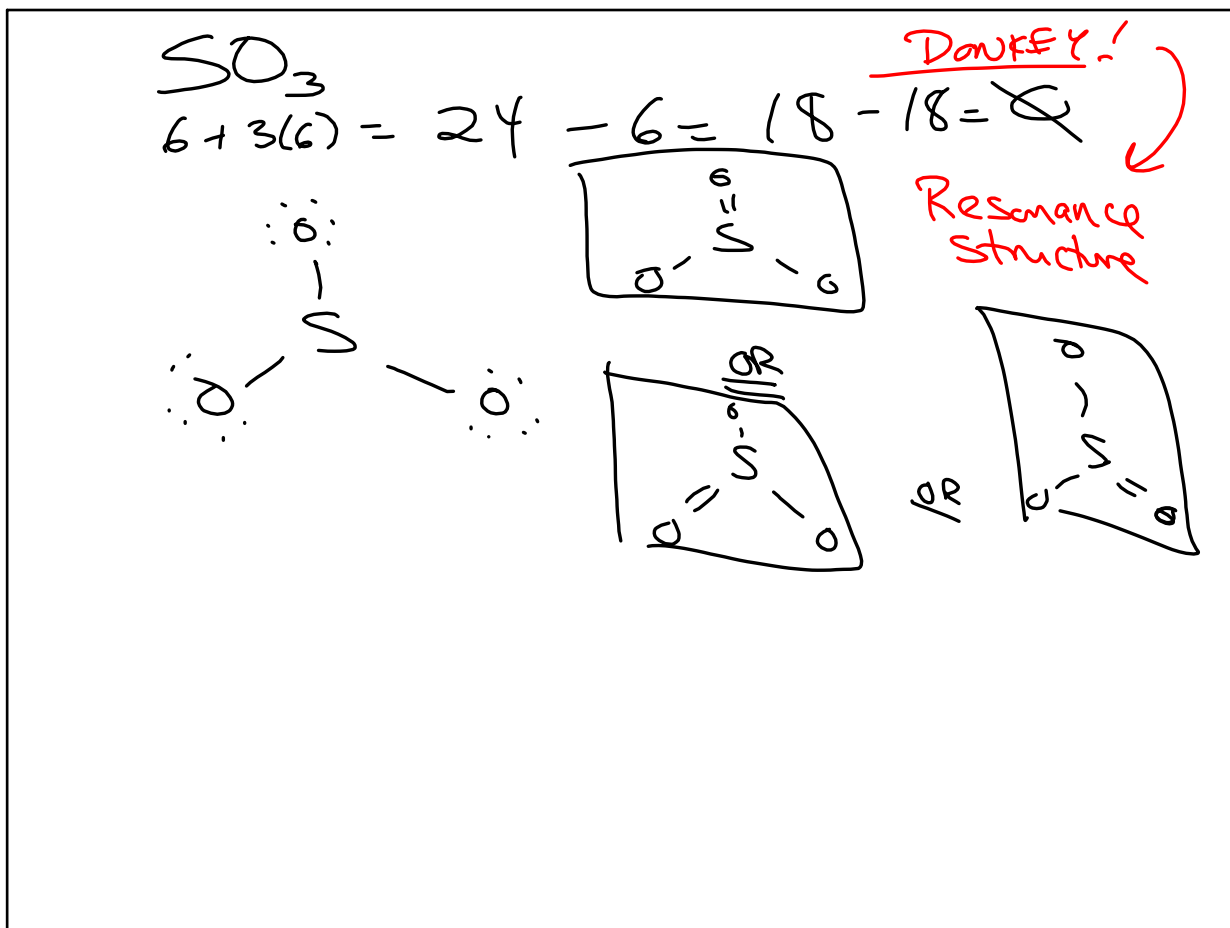
Nov 12-9:09 AM



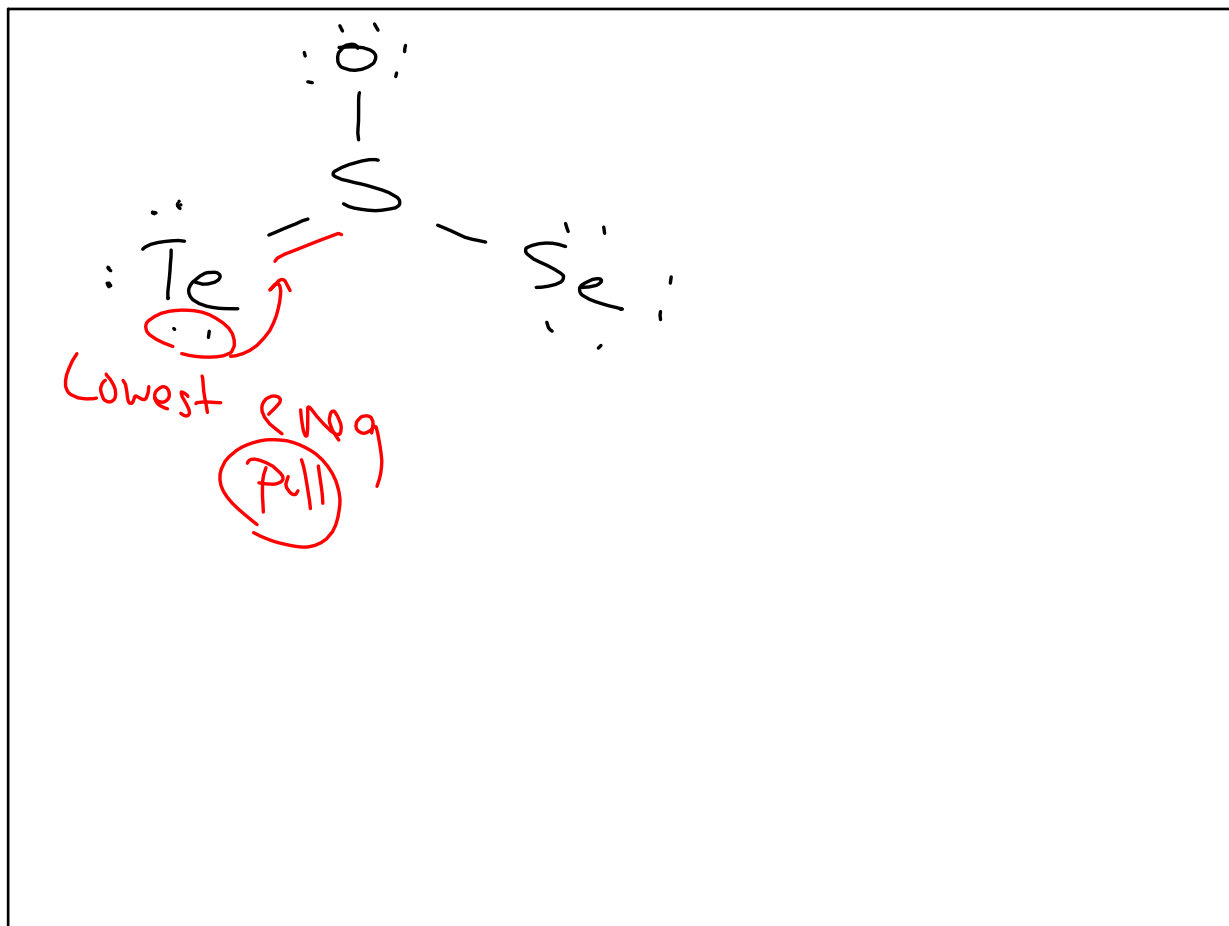
Nov 12-9:16 AM



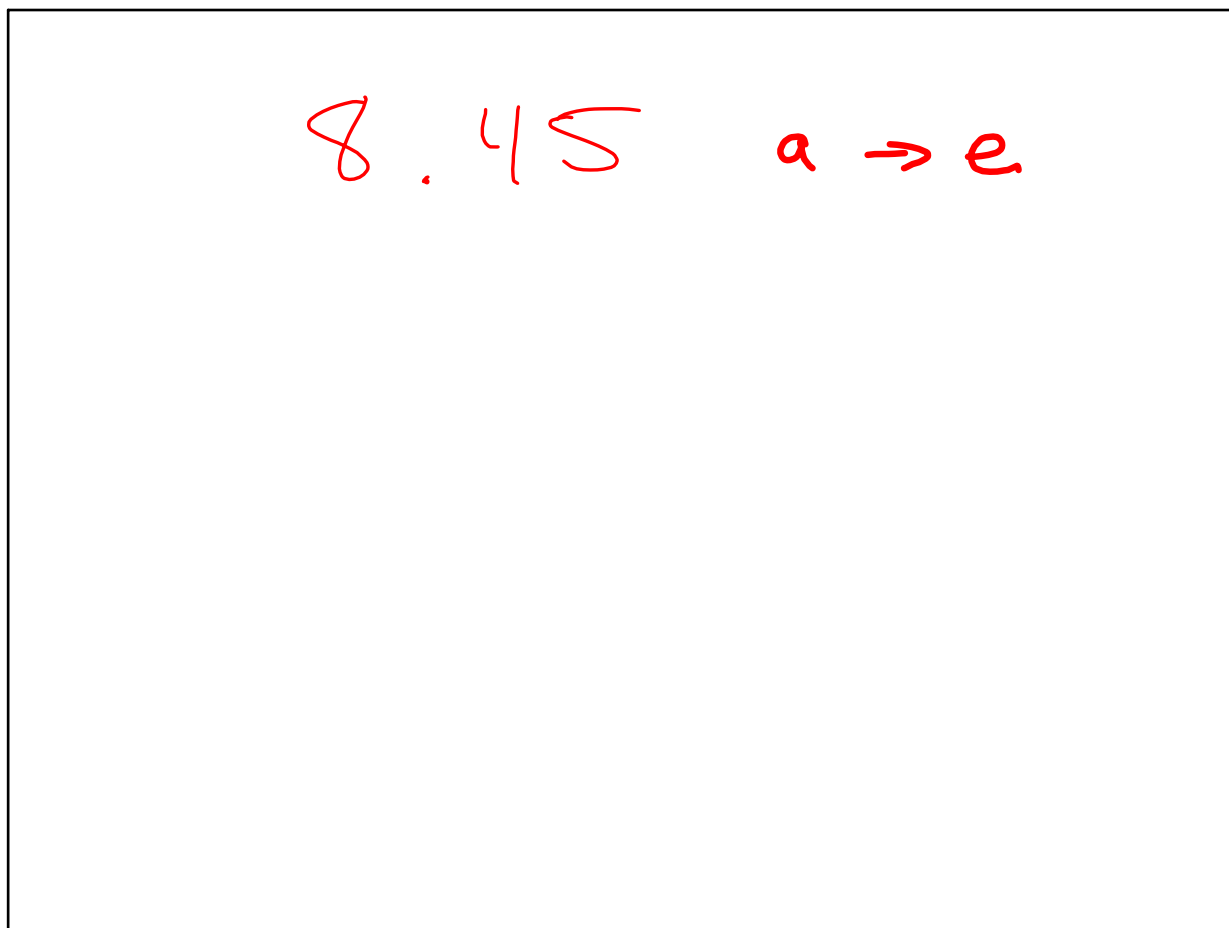
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Nov 12-9:21 AM



Nov 12-9:24 AM



Nov 12-9:30 AM