

PS 3-1

60% X

⑨ X_2O_3

$X = \frac{50g}{mole}$

$O = \frac{?g}{mole}$

$\frac{X_2}{X_2O_3} = \frac{0.6}{1}$

$\frac{100}{X_2O_3} = \frac{0.6}{1}$

$\frac{100}{0.6} = X_2O_3$

$166.67g = X_2O_3$

$\frac{O_3}{X_2O_3} = 0.4$

$X_2O_3 = 166.67$

$2(X) + 3(O) = 166.67$

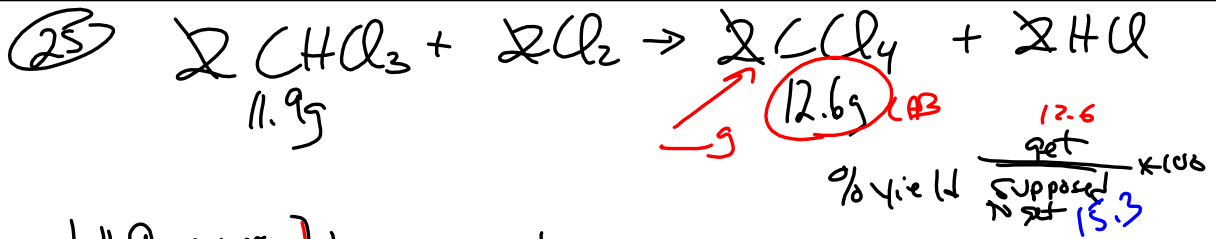
$3(O) = 66.67$

$O = 22.22 g/mole$

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5g	12
	13

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11.9g CHCl₃	1 mole CHCl ₃	1 mole CCl ₄	152g CCl ₄ = 15.3
11.9g CHCl ₃	1 mole CHCl ₃	1 mole CCl ₄	

$$\frac{12.6}{15.3} \times 100 = 82.19\%$$

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Chap 4 - Solution Chem.

Solute dissolved in a solvent to produce a homogeneous mixture.

- S + l ← salt/water
- l + l ← dilute
- g + l ← SODA (Pop)
- S + S ← Bronze / Brass
- ↙ Alloy
- g + g ← Air

Guyana
Tim Jones
☹ Poison

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$S+l \rightarrow$ always a soln?
 $l+l \rightarrow$ NO!

$Oil + H_2O$
 Non-Polar \rightarrow POLAR

Oil + vinegar
 Italian dressing

$H_2O \rightarrow$ Polar + Formic
 Not Non-Polar

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3.77

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