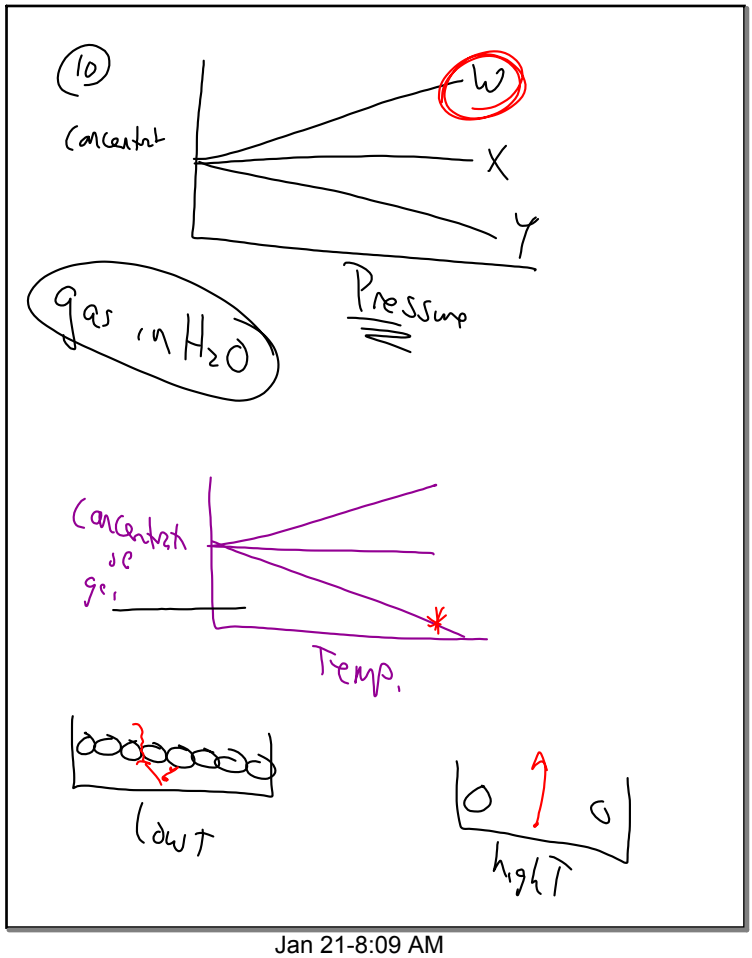

Jan 21-8:09 AM

Ex 1 15g NH_3 in 250g H_2O, $D_{soln} = \frac{0.974\ g}{ml}$

I Mole fraction NH_3 $= \frac{\text{Mole } NH_3}{\text{Mole } NH_3 + \text{mole } H_2O}$ $\left(\frac{g}{MW}\right)$

$= \frac{\frac{15}{17}}{\frac{15}{17} + \frac{250}{18}} = 0.06$

II Molality $NH_3 = \frac{\text{Moles } NH_3}{\text{Kg } H_2O} = \frac{\frac{15}{17}\ (\text{mole})}{0.250\ kg} = 3.529\ m \approx 3.53$

III $M = \frac{\text{Mole } NH_3}{\text{l of solution}} = \frac{\frac{15}{17}}{0.272\ l} = 3.26\ M$

1ml	265g
0.974g soln	$= 272.09$ mp

Solute $15g\ NH_3$ Solvent $250g\ H_2O$

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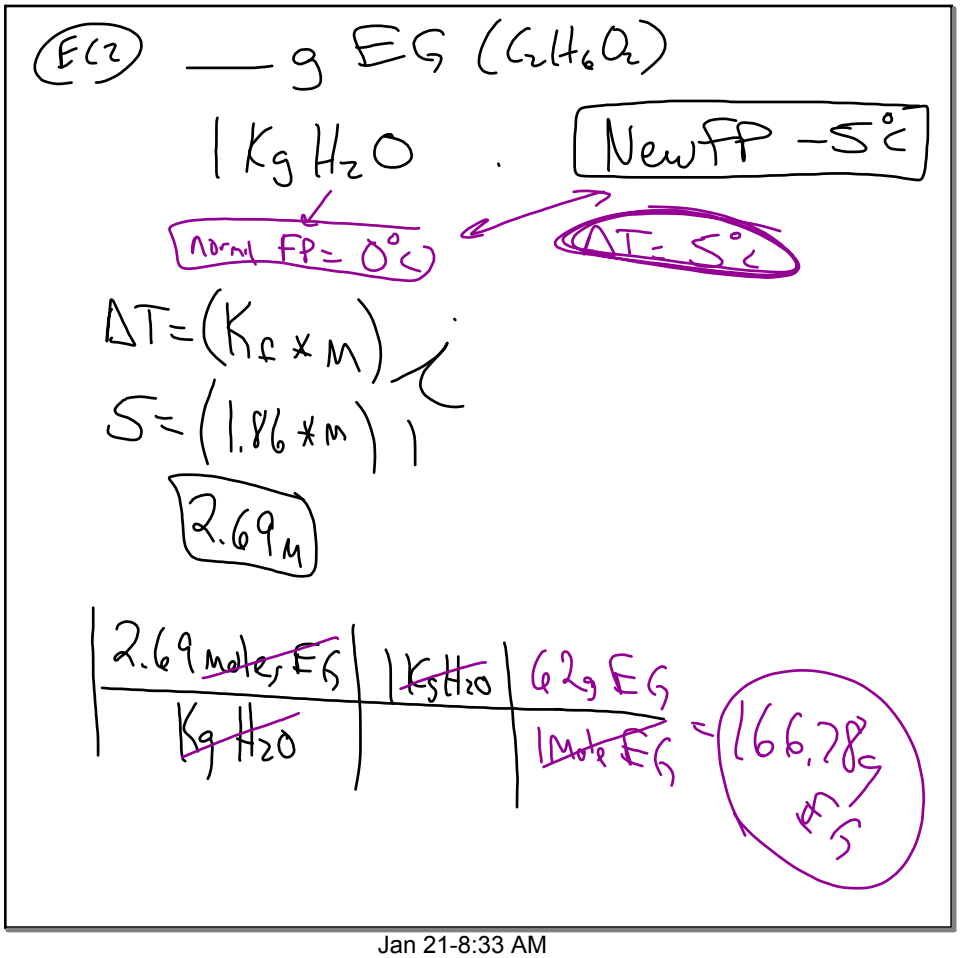

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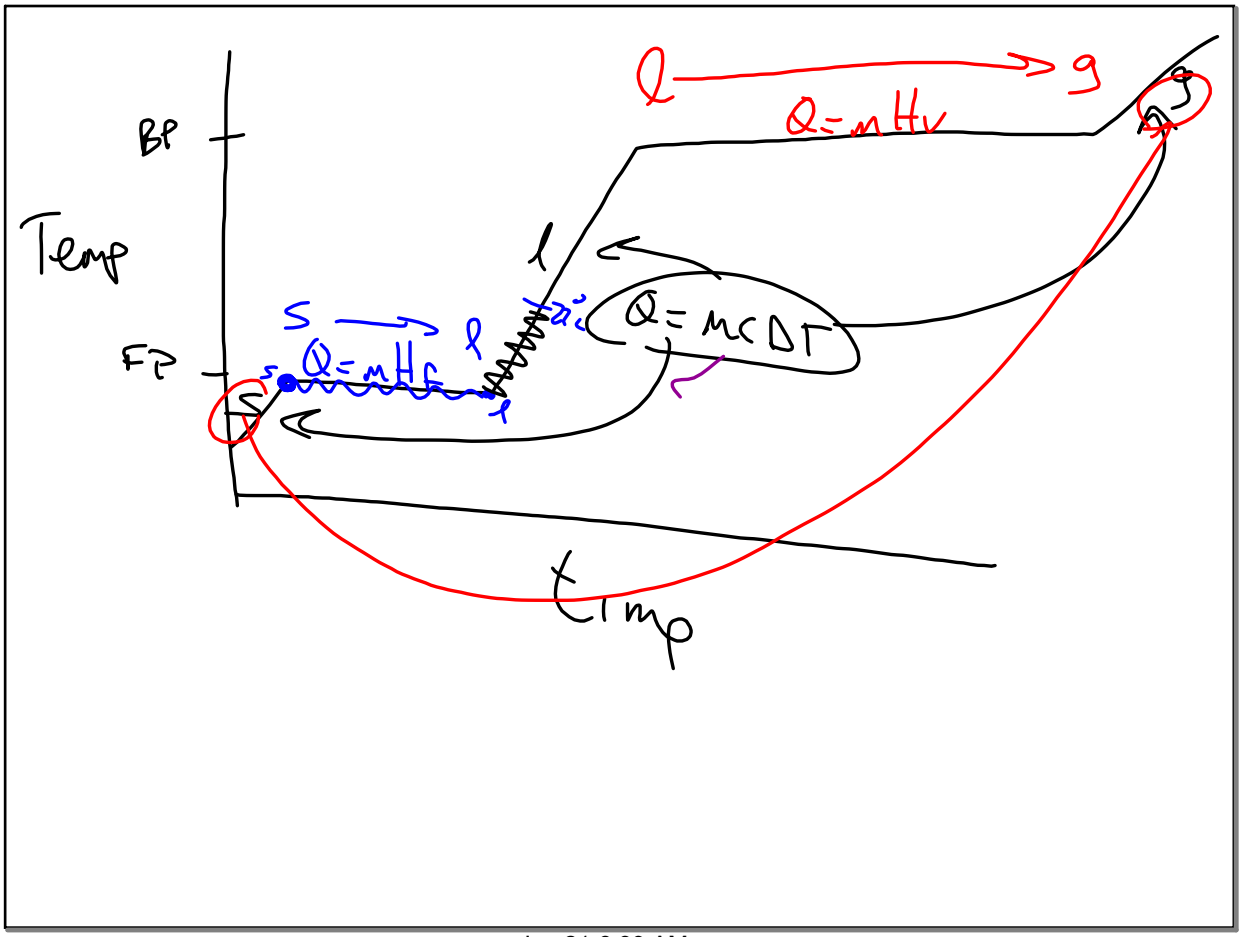
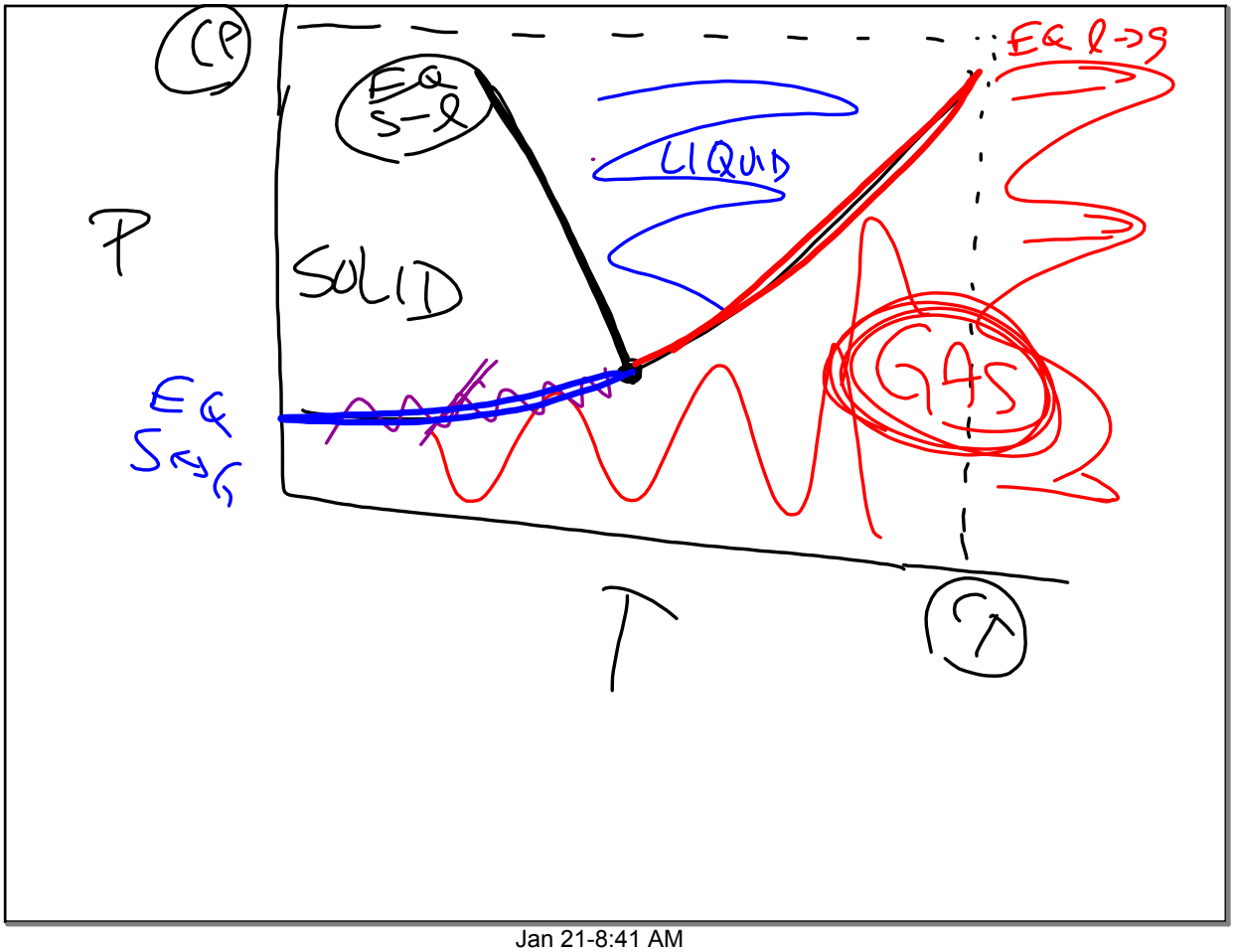
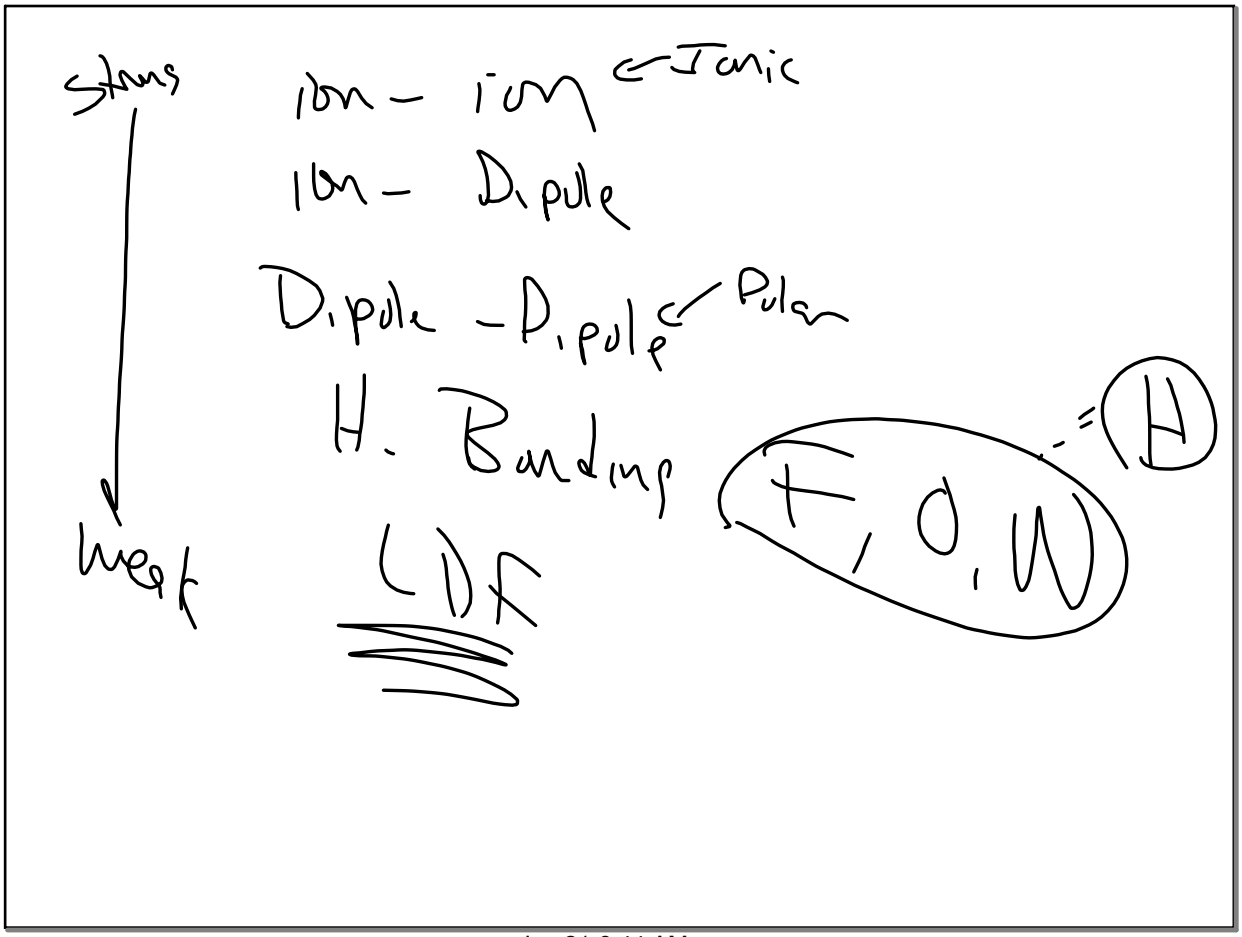


Jan 21-8:39 AM



Jan 21-8:41 AM



Jan 21-8:44 AM