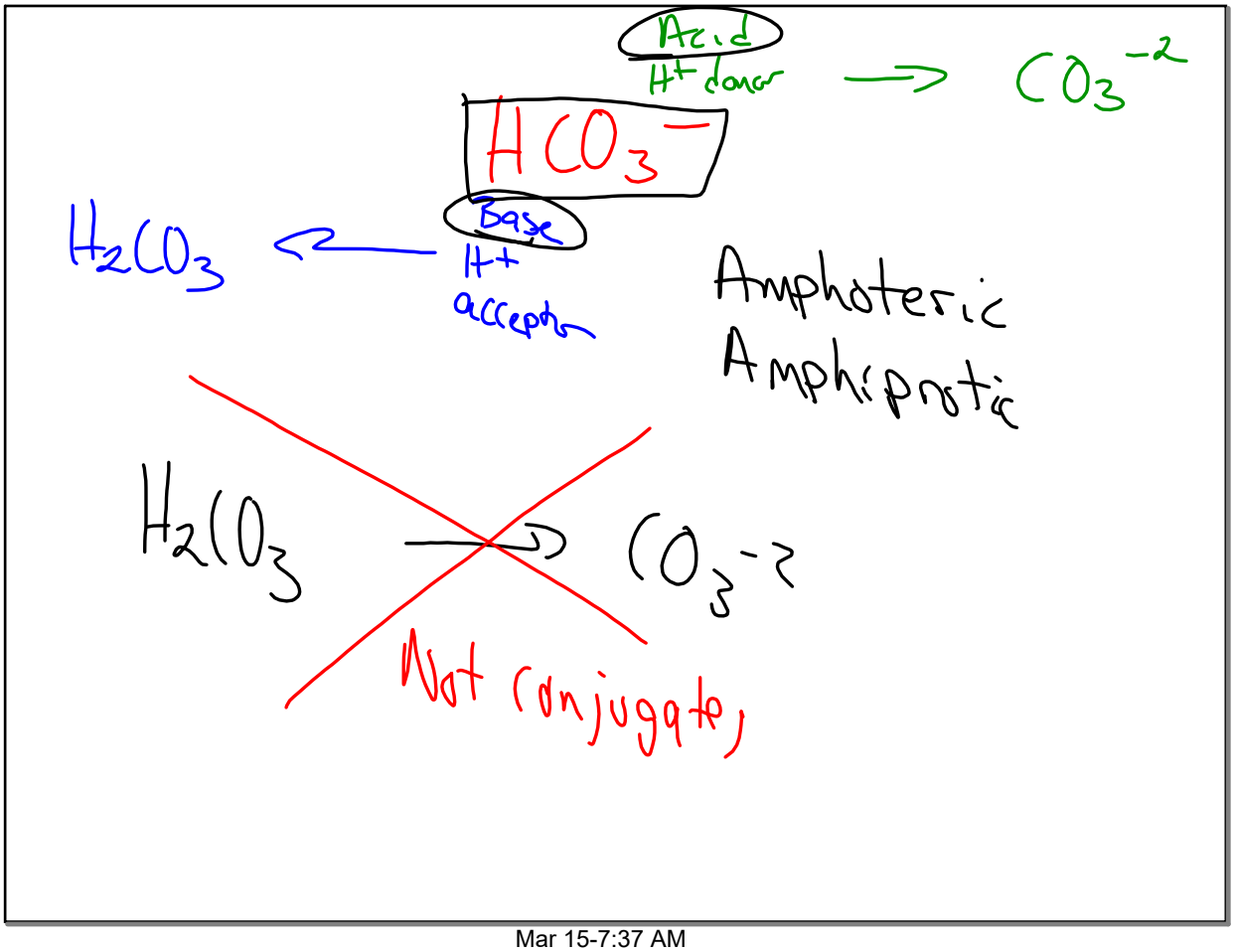
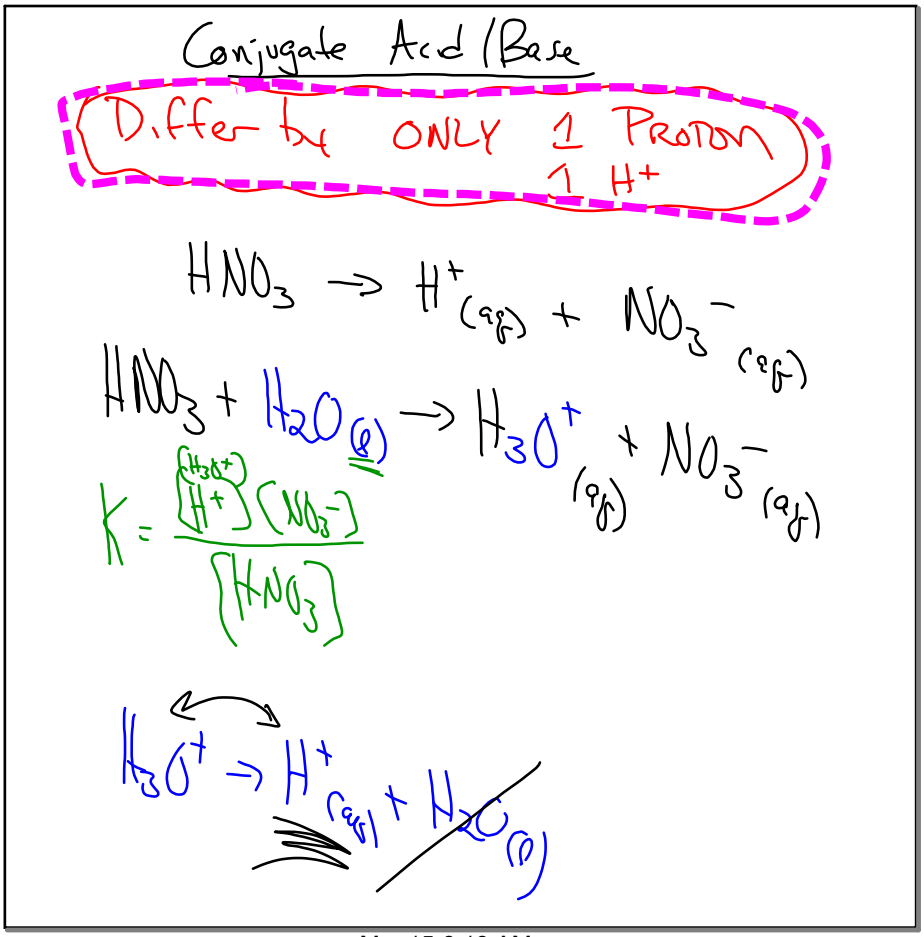


Mar 15-7:37 AM



Mar 15-8:18 AM

Low pH Acid Strength → More H^+ ions
"Below 7"
$[H^+] > [OH^-]$

High pH Base Strength → More OH^- ions.
>7
$[OH^-] > [H^+]$

Power of H^+ ⇒ pH

Mar 15-8:24 AM

$pH = -\log[H^+]$

$pOH = -\log[OH^-]$

$pH + pOH = 14$

$K_w = [H^+][OH^-] = 1 \times 10^{-14}$ at 25°C

H_2O
$HOH(l) \rightarrow H^+_{(aq)} + OH^-_{(aq)}$

Mar 15-8:26 AM

16.22 $H_2C_6H_7O_5^- + H_2O(l) \rightarrow$

BASE

Conj pairs

ACID

$H_3C_6H_7O_5$ + OH^-

Mar 15-8:37 AM

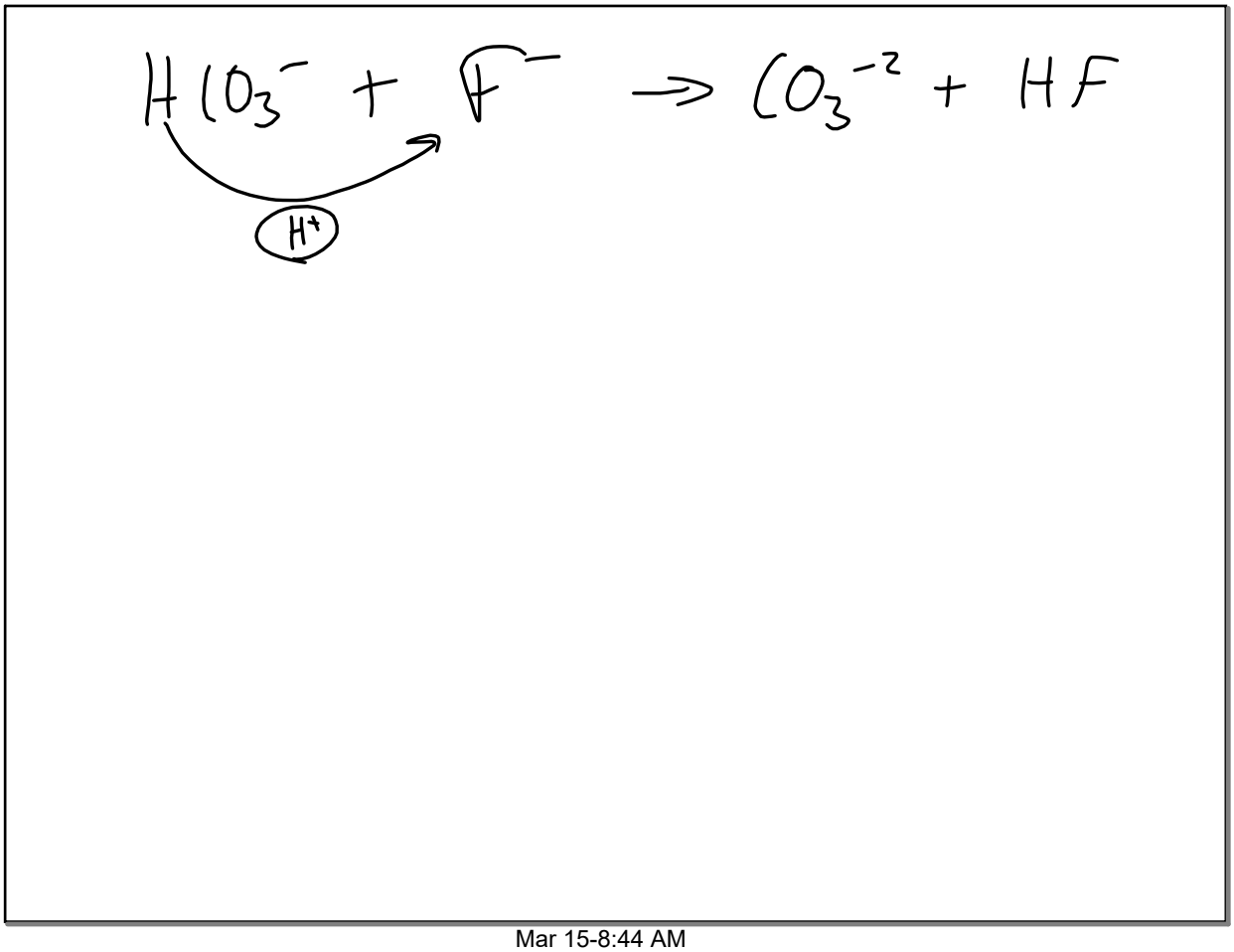
$CH_3COO^- + H_3O^+ \rightarrow$

H^+

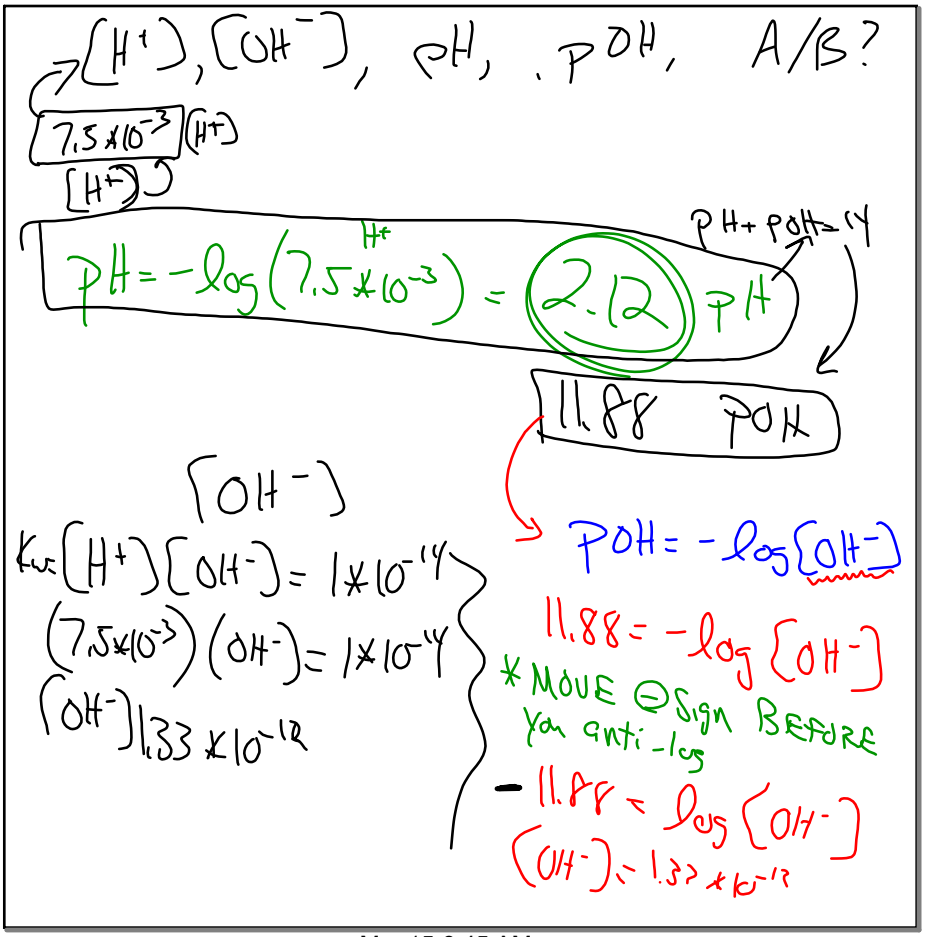
CH_3COOH + H_2O

ORG ACID

Mar 15-8:42 AM



Mar 15-8:44 AM



Mar 15-8:45 AM

$pH = -\log[H^+]$

$pOH = -\log[OH^-]$

$K_w = [H^+][OH^-] = 1 \times 10^{-14}$ at 25°C

$pH + pOH = 14$

Mar 15-8:59 AM

Strong Acids

Dissociate (Break up) ≈ 100% (into ions)

	$1\ HNO_3(aq)$	→	$1\ H^+(aq)$	+	$1\ NO_3^-(aq)$
R					
I	2M		0		0
C (Δ)	−2		+2		+2
E	≈ 0M		2M		2M

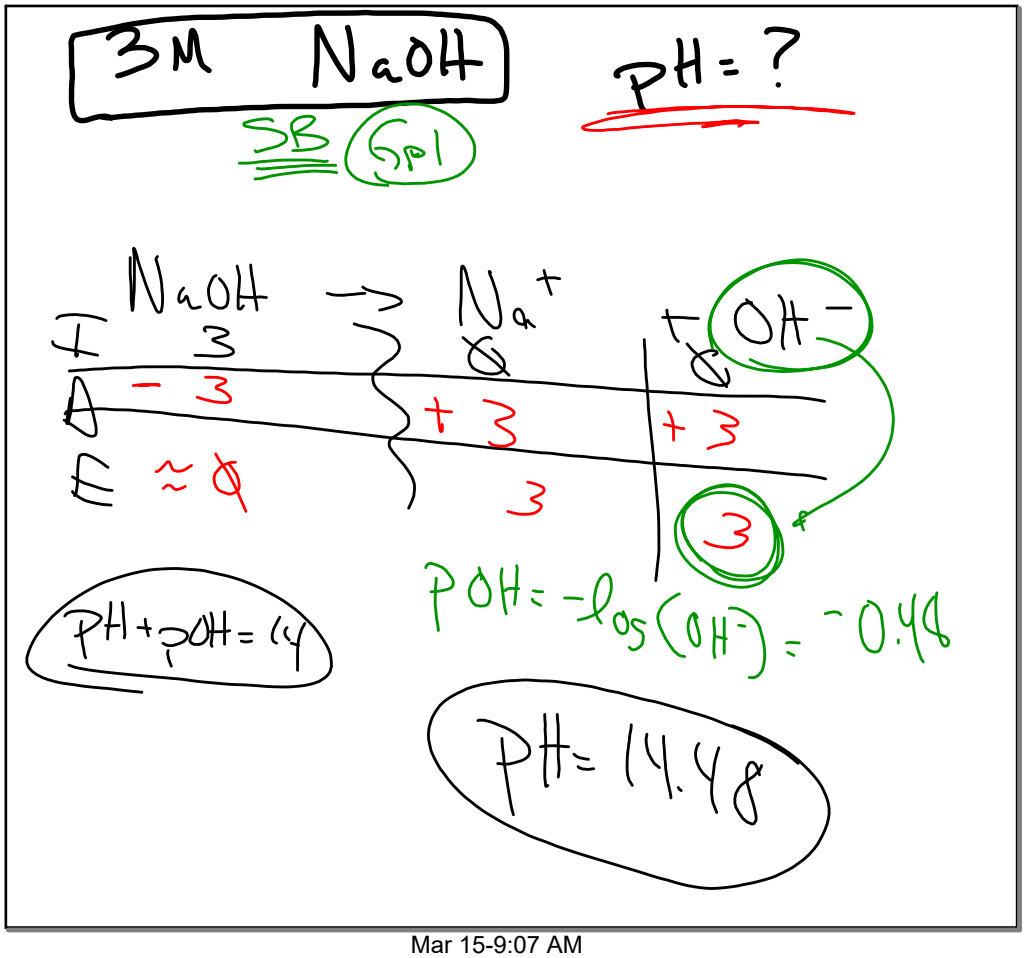
MOLE RATIO

$pH = -\log[H^+]$

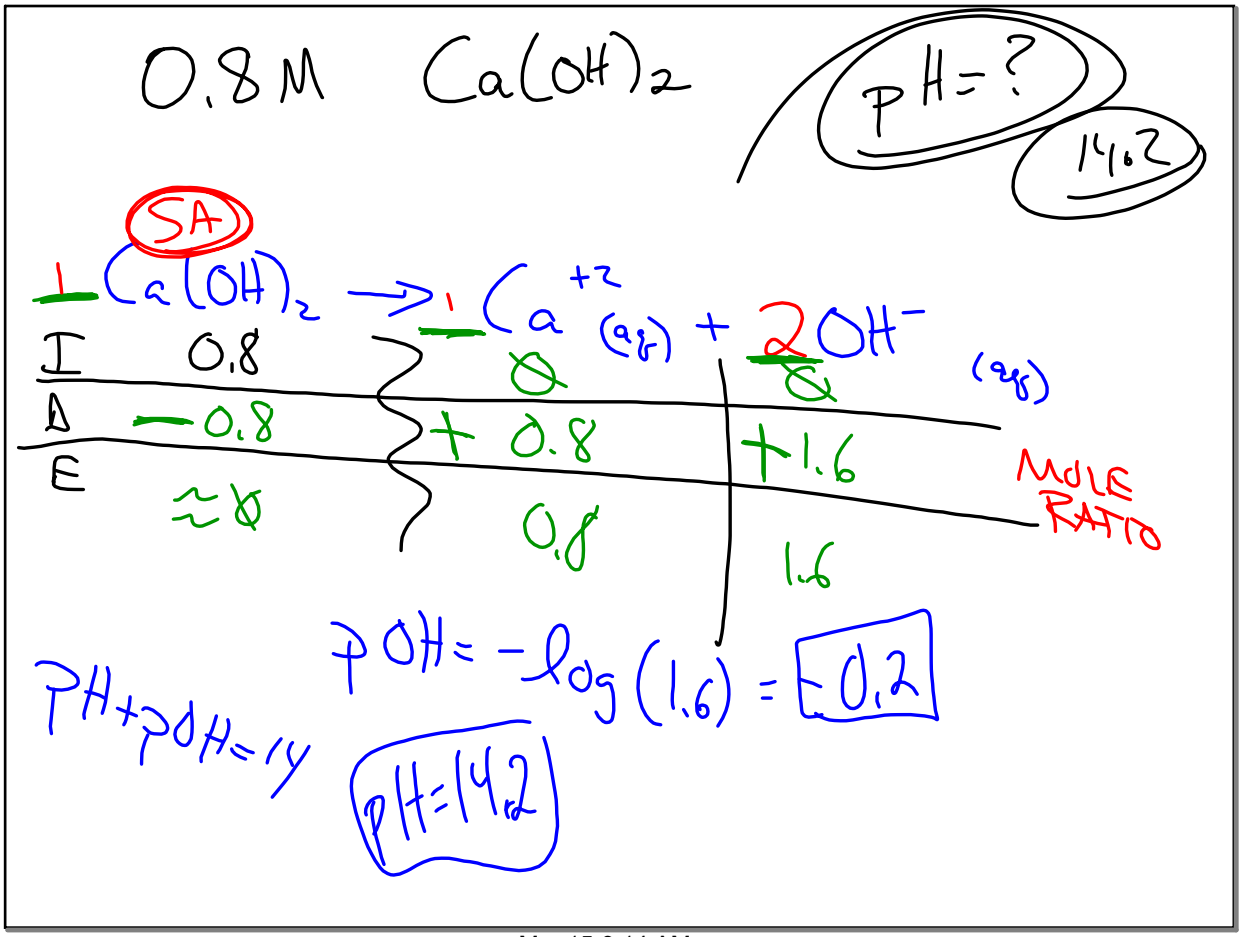
$pH = -0.3$

$pOH = 14.3$

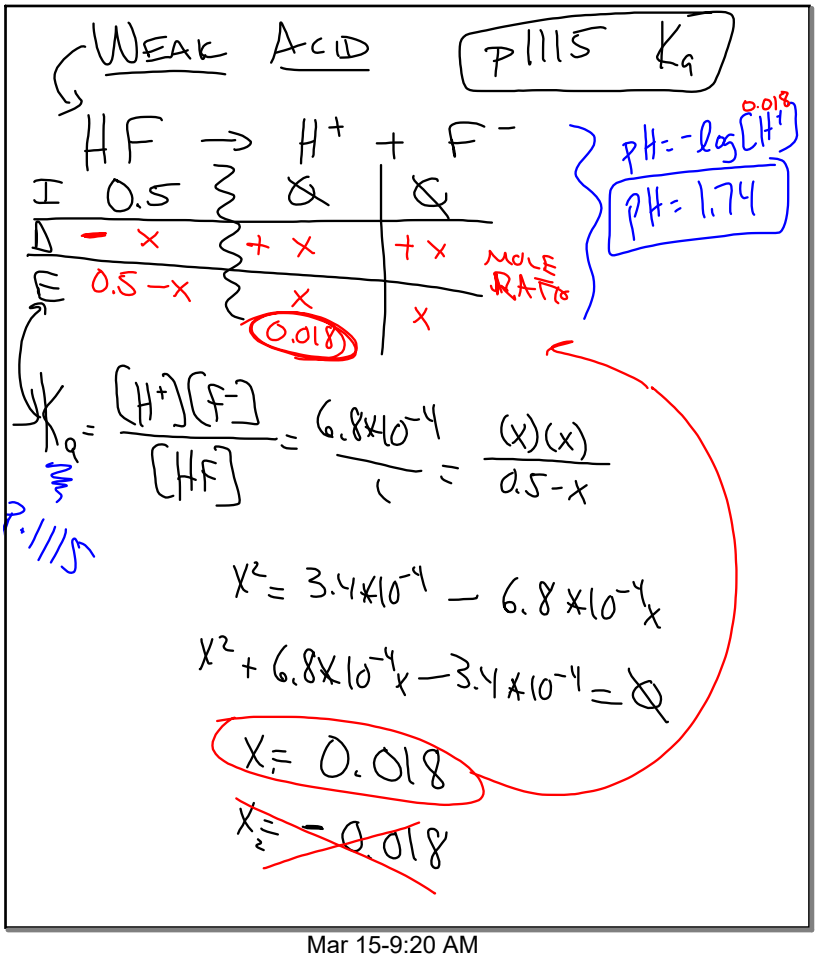
Mar 15-9:00 AM



Mar 15-9:07 AM



Mar 15-9:14 AM



Mar 15-9:20 AM

HW

16/ 49, 53, 57, 63

RICE

Mar 15-9:31 AM