

Solvent dissolves the solute \Rightarrow Solution

Present in greater quantity

Hydration

Solvation

Water "tears" the crystal apart.
Overcome the crystal lattice structure (Energy)

Jan 12-7:20 AM

Types of Solutions

<u>Unsaturated</u>	<u>Saturated</u>	<u>SUPER Saturated</u>
<p style="color: red;"><u>"VACANCY"</u></p> <p>Solvent <u>can</u> hold more solute</p> <p>Dilute "Watery" little solute</p> <p>Concentrated lots of solute</p>	<p style="color: red;"><u>"NO VACANCY"</u></p> <p style="color: red;"><u>Full</u></p> <p>The solvent is completely saturated (full) with solute.</p> <p style="color: red;"><u>Bus Full</u></p>	<p style="color: green;"><u>"Special conditions"</u></p> <p>That allow the sat'd solvent to temporarily hold more solute</p>

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Measure how much solute is in the solvent. → How concentrated.

(M) Molarity = $\frac{\text{Moles of solute}}{\text{l of solution}}$

↓ ↓
Solute + Solvent

(m) molality = $\frac{\text{Moles of solute}}{\text{Kg of Solvent}}$

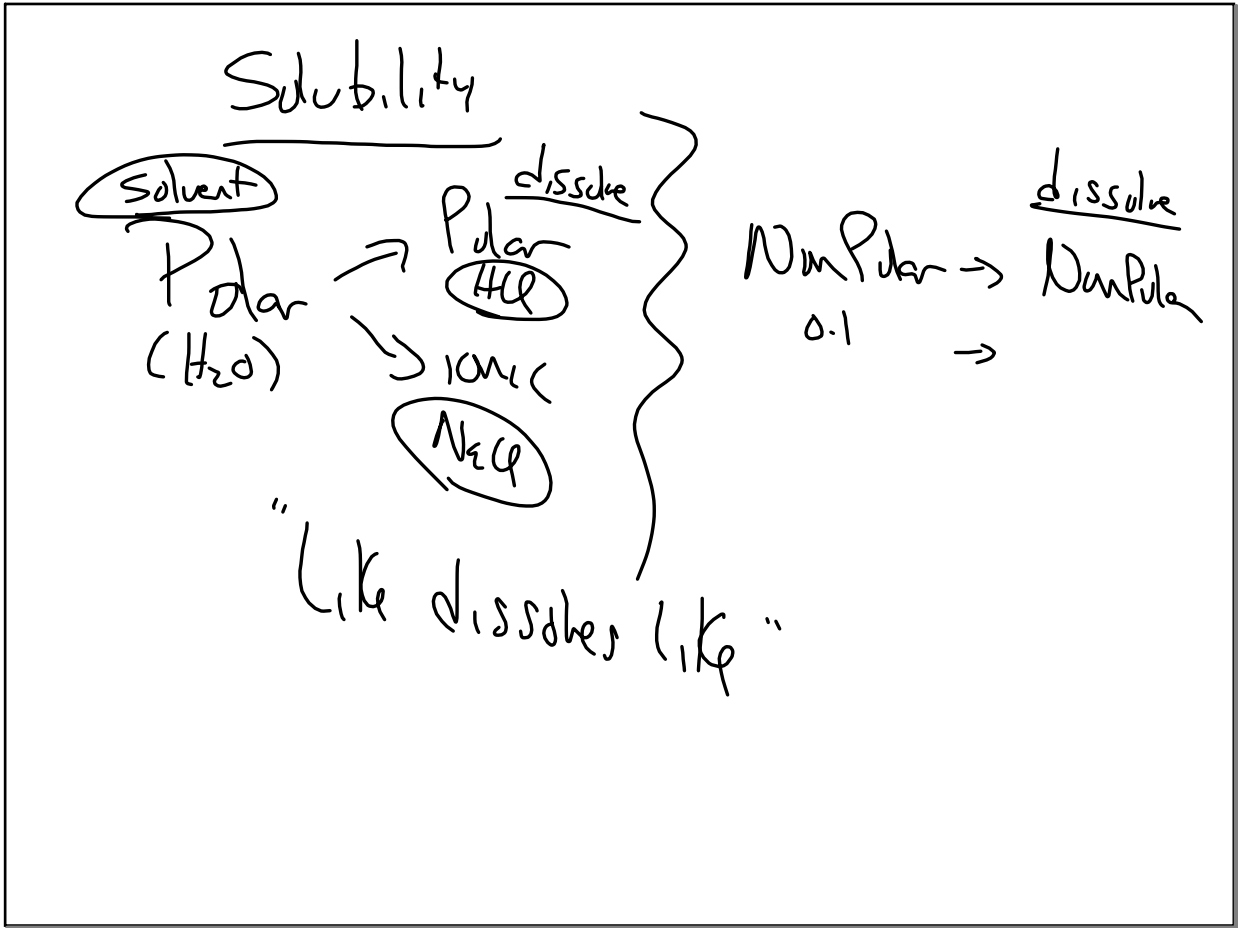
$D = \frac{m}{V}$

Jan 12-8:40 AM

% composition ^{pph} or PPM.

$\frac{\text{Part}}{\text{Whole}} \times 100$ } $\frac{\text{Part}}{\text{Whole}} \times 1,000,000$

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