

7125 Solubility Chart

Spectator ion \rightarrow comes to watch
The game and leaves \rightarrow No change.

Chem \rightarrow comes + goes without a change
 $Cl^{-}(aq) \rightarrow Cl^{-}(aq)$

Part of reaction

Change $(l) \rightarrow (s)$

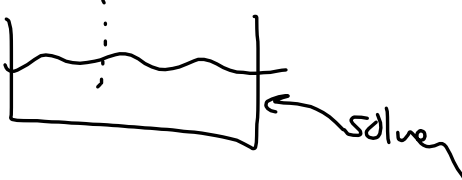
Oct 6-8:05 AM

4/7 (9) Cl^{-} spectator
except Ag^{+} , Hg_2^{+2} , Pb^{-2}

(6) NO_3^{-}

Oct 6-8:22 AM

Solute



← solvent

Dissolves evenly

↓

Solution

$$\frac{\uparrow \text{Amount of solute}}{\text{Solution}} = \boxed{\frac{\text{Moles solute}}{\text{l of solution}}}$$

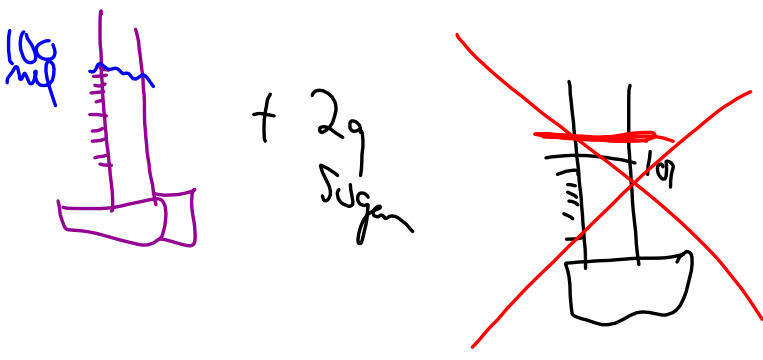
Measure Concentration

Molarity

Oct 6-8:27 AM

$$\frac{M}{l} = \frac{\text{Moles of solute}}{\text{l of solution}}$$

Solute + solvent.

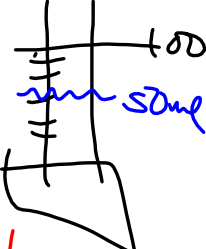


100 ml

+ 2g Sugar

Oct 6-8:37 AM

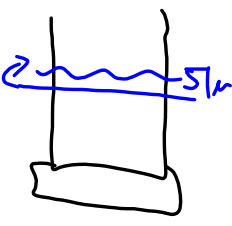
100ml solution



less solvent

+2g Sugar

Add Solute



Add H₂O to 100ml

Mix Solution

Oct 6-8:38 AM

2 Molar solution HCl

2 moles HCl

1 L

F L M

Oct 6-8:40 AM

Prepare 250ml 1M HCl

FCM - unit to find gues ON TOP

250ml \rightarrow 0.250 L

1 mole HCl
1 L HCl

1 mole HCl	0.250 L	36g HCl
1 L HCl		1 mole HCl

g ?

Oct 6-8:40 AM

HW 4/62

Oct 6-8:44 AM