

PS 3-1

③ 50g NaCl + 500g H<sub>2</sub>O

$$\% \text{ by mass NaCl} = \frac{\text{Part}}{\text{WHOLE}} = \frac{\text{NaCl}}{\text{NaCl} + \text{H}_2\text{O}}$$

$$= \frac{50}{50 + 500} \times 100$$

$$= 9.1\%$$

Oct 13-8:22 AM

⑤ — g H<sub>3</sub>PO<sub>4</sub>, 175ml, ~~3.5M H<sub>3</sub>PO<sub>4</sub>~~

↓

0.175 l

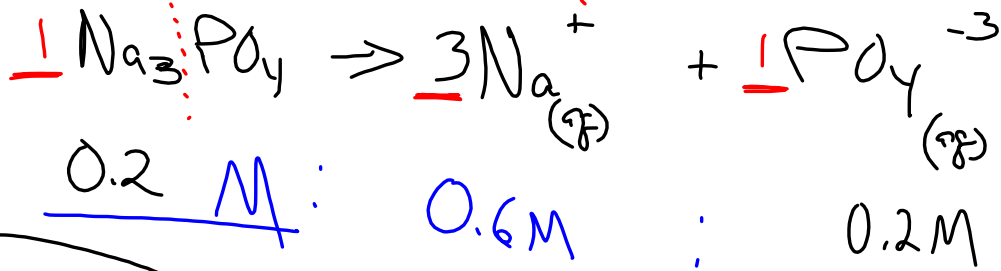
3.5 mols H<sub>3</sub>PO<sub>4</sub>  
1 l

<del>3.5 mols H<sub>3</sub>PO<sub>4</sub></del> 1 l	0.175 l	<sup>3(1)+31+4(16)</sup> 98 g H <sub>3</sub> PO <sub>4</sub> <del>1 mol H<sub>3</sub>PO<sub>4</sub></del>	= 60 g H <sub>3</sub> PO <sub>4</sub>
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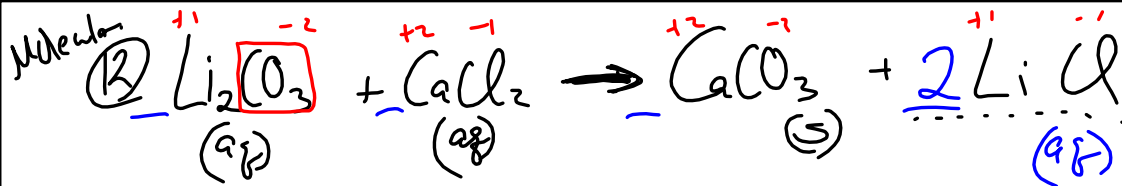
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⑦ ? M  $\text{Na}_3\text{PO}_4$ , Given:  $0.6\text{M Na}^+$

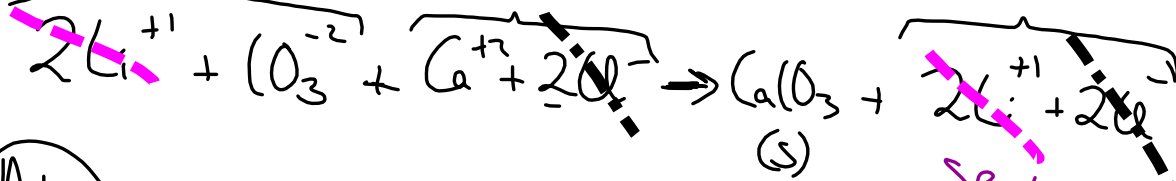


MOLE RATIO

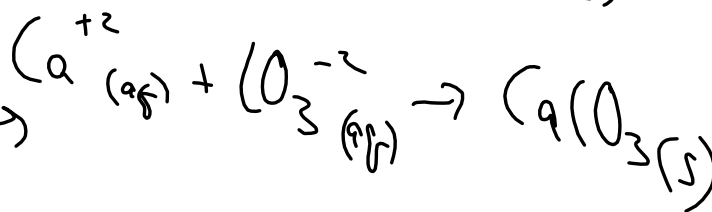
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(complete ion)



Net



Spectator ions

Oct 13-8:37 AM