Formula mass / GFM / mole $C_6H_{12}O_6$

6 C = 6(12)

12 H = 12(1)

6 O = 6(16)

140 grams / Mole = GFM

Sep 25-8:06 AM

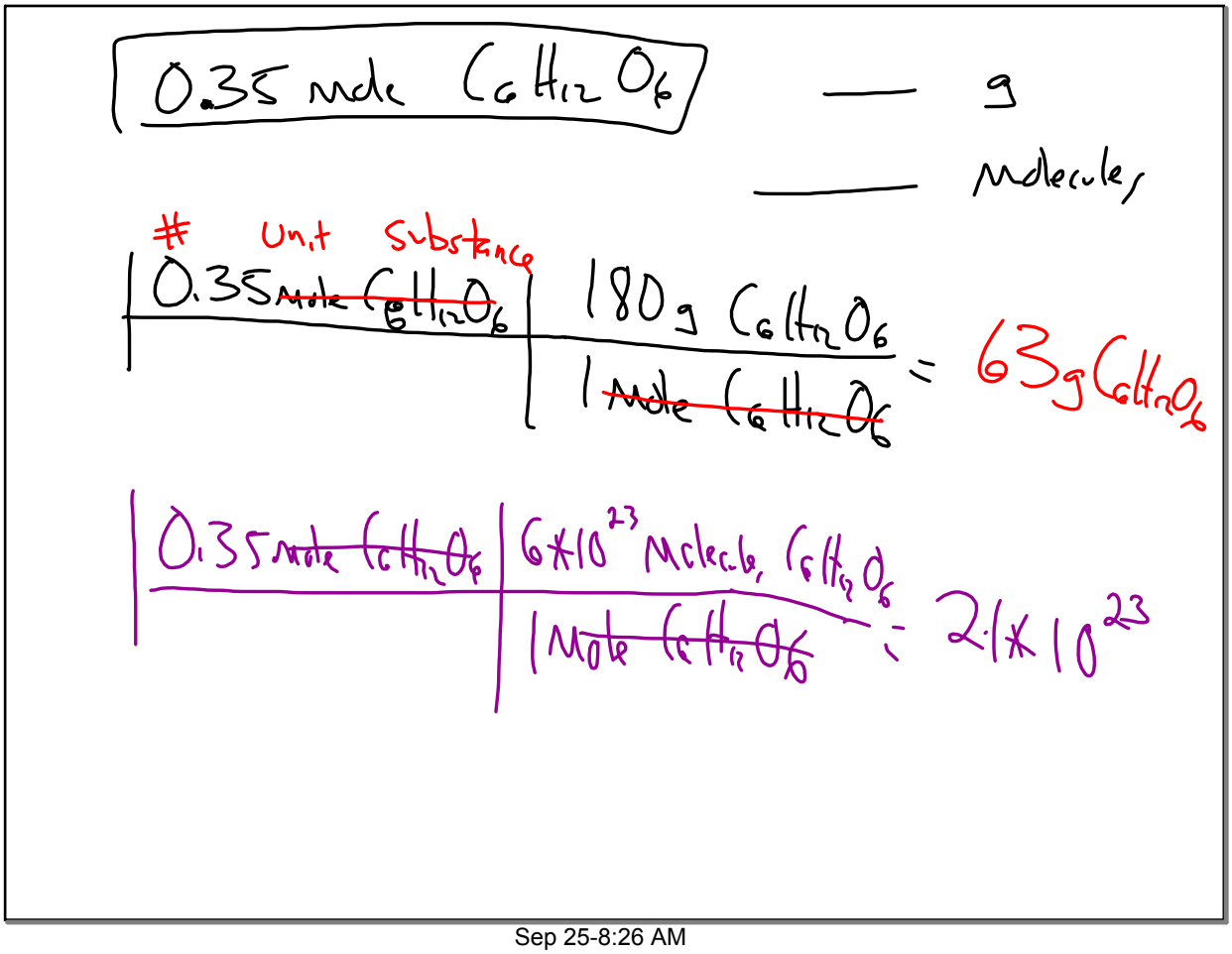
% Composition $\frac{\text{Part}}{\text{Whole}} * 100$

$C_6H_{12}O_6$

% C by mass in $C_6H_{12}O_6$ $\frac{\text{Part}}{\text{Whole}} = \frac{6\,C}{C_6H_{12}O_6} = \frac{6(12)}{180}$

$\frac{72}{180} * 100$ 40% by mass

Sep 25-8:24 AM

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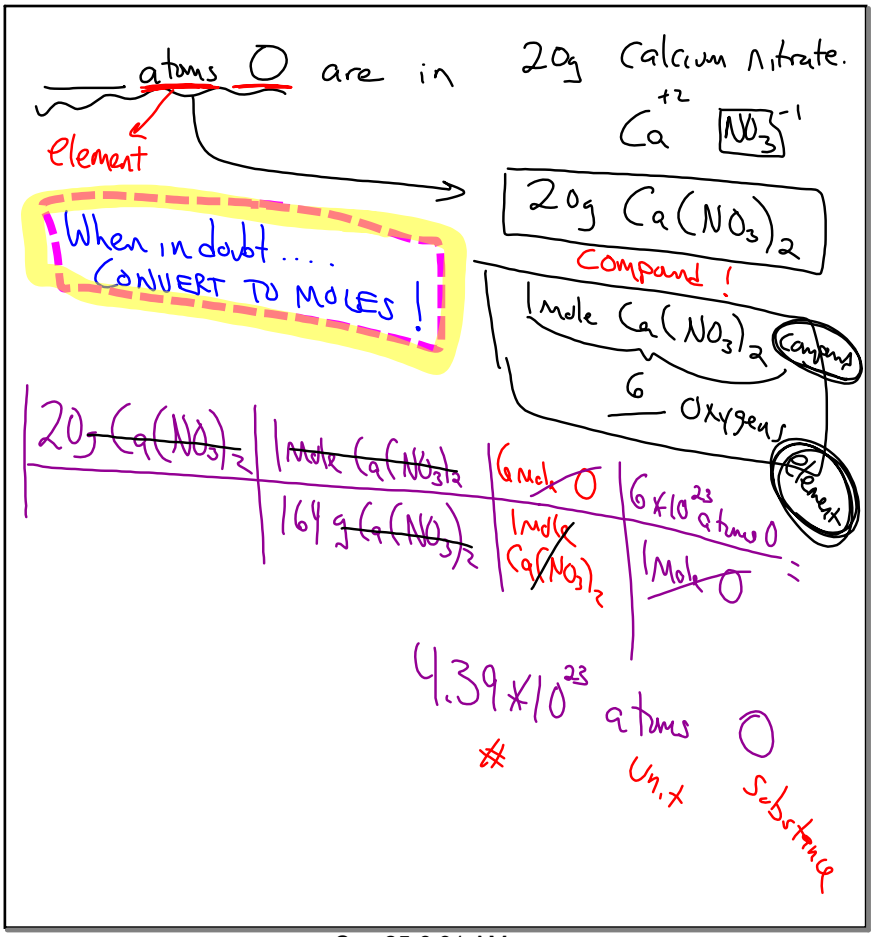


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3/ 22 a+b

24 a+b

34 all.

S. A. W

Show ALL WORK

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