

$2Al(s) + 6HCl(aq) \rightarrow 2AlCl_3(s) + 3H_2(g)$

$< 0.041g$ (red)
 3ml 4M (blue)
 0.035g Al (black)

④ Mole Al : Mole H₂
 Exp

Actual Al H₂
 2 : 3

% error

PV = nRT
 n = PV / RT
 PV ← syringe
 RT ← temp

① ml(g) + 1.2ml
 "Dome of syringe"

② P_T = P_{H₂} + P_{H₂O}
 P_{H₂} DRY
 P_{atm}
 T_{H₂O} (chart)

③ Find mole H₂
 Mole H₂ (circled)
 2 mole Al (circled)
 3 mole H₂ (circled)
 Mole Al (circled)

⑤ $\frac{g}{mole} = \frac{Al \text{ used to start}}{mole Al}$ → molar mass (MW)

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2014
Final # 1-10

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