

Enthalpy of a Reaction ΔH_{rxn}

$$\Delta H_{rxn} = \Delta H_{prod} - \Delta H_{reactants}$$

$\oplus \Delta H$

Endothermic
Non-spont. ;

$\ominus \Delta H$

Exothermic
SPONTANEOUS

(P - R)

Oct 21-7:38 AM

2 : 1 : 2 ; -483.6 kJ

2 H₂ (g) + O₂ (g) → 2 H₂ O ΔH = -483.6 kJ

ΔH IS IS Part of the mole RATIO

lose energy
SPONT ;

5g H₂	1 mole H₂	-483.6 kJ	=	() kJ
2g H ₂	2 mole H ₂	-483.6 kJ		

Oct 21-7:50 AM

HW
S/421

Oct 21-7:55 AM