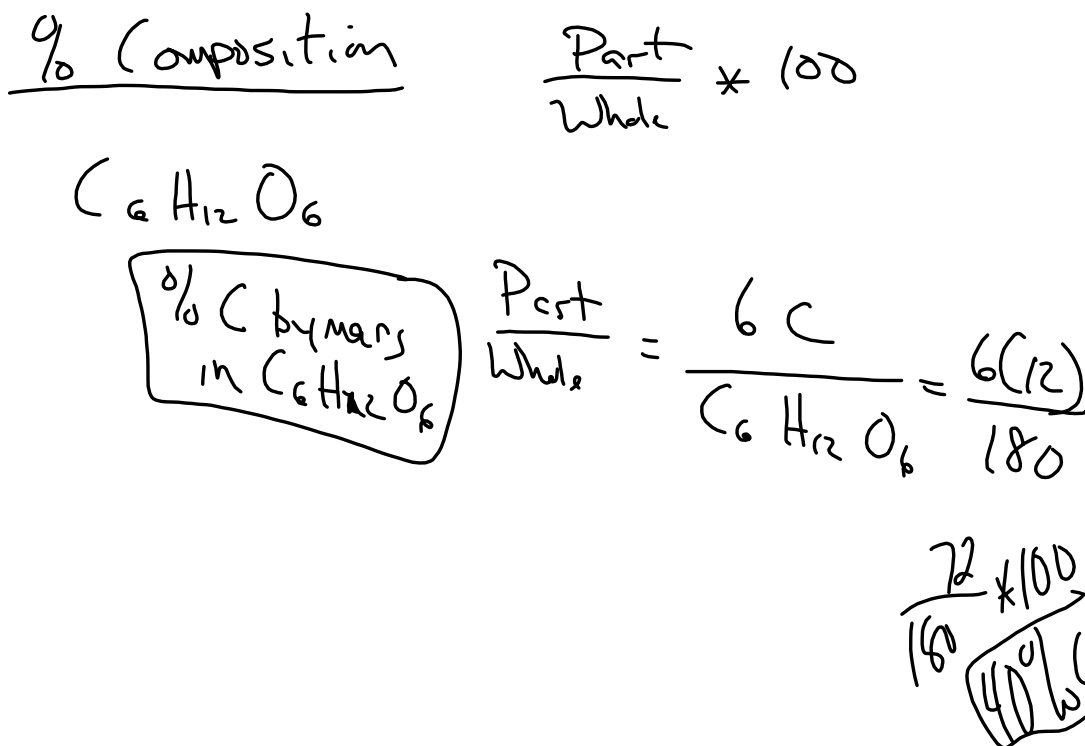


Sep 25-8:06 AM



Sep 25-8:24 AM

0.35 mole CaH_2O_6 — g
 — molecule

Unit Substance
 $\frac{0.35 \text{ mole } \text{CaH}_2\text{O}_6}{1 \text{ mole } \text{CaH}_2\text{O}_6} \times 180 \text{ g } \text{CaH}_2\text{O}_6 = 63 \text{ g } \text{CaH}_2\text{O}_6$

$\frac{0.35 \text{ mole } \text{CaH}_2\text{O}_6}{1 \text{ mole } \text{CaH}_2\text{O}_6} \times 6 \times 10^{23} \text{ Molecules } \text{CaH}_2\text{O}_6 = 2.1 \times 10^{23}$

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atoms O are in 20g Calcium Nitrate.
 element $\text{Ca}^{+2} \text{NO}_3^{-1}$

When in doubt...
 CONVERT TO MOLES!

20g $\text{Ca}(\text{NO}_3)_2$ Compound!
 1 mole $\text{Ca}(\text{NO}_3)_2$ Compound
 6 Oxygen Element

$\frac{20 \text{ g } \text{Ca}(\text{NO}_3)_2}{164 \text{ g } \text{Ca}(\text{NO}_3)_2} \times \frac{6 \text{ mole O}}{1 \text{ mole } \text{Ca}(\text{NO}_3)_2} \times 6 \times 10^{23} \text{ atoms O} = 4.39 \times 10^{23} \text{ atoms O}$

Unit Substance

Sep 25-8:31 AM

3 / 22 a+b
24 a+b
34 all.

S. A. W
Show All Work

Sep 25-8:44 AM